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## **Modern Chemistry Chapter 3 Mixed**

CHAPTER 3 REVIEW  
Atoms: The Building  
Blocks of Matter MIXED  
REVIEW SHORT  
ANSWER Answer the  
following questions in  
the space provided. 1.

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## Chapter 3 Mixed Review Answers

The element boron, B, has an atomic mass of 10.81 amu according to the periodic table. However, no single atom of boron has a mass of exactly 10.81 amu. How can you explain this difference?

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Chapter 3. law of conservation of mass.  
atom. nucleus. proton.  
states that matter is neither created nor destroyed during ord....  
the smallest particle of

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an element that retains  
the chemical... is a  
very small region  
located at the center of  
an atom.

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CHAPTER 3 REVIEW.  
Atoms: The Building  
Blocks of Matter.

MIXED REVIEW. SHORT  
ANSWER Answer the  
following questions in  
the space provided. 3

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## Chapter 3 Mixed Review Problems

points each, 9 points total.

1. The element boron, B, has an atomic mass of 10.81 amu according to the periodic table. However, no single atom of boron has a mass of exactly 10.81 amu. How can you explain this difference?

2.

### **CHAPTER 3 REVIEW - Doral Academy Preparatory School**

3 at 25°C is 10. g of



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## Chapter 3 Mixed Review Answers

solute per 100. g of H<sub>2</sub>O. a. If 15 g of KClO<sub>3</sub> are stirred into 100 g of water at 25°C, how much of the KClO<sub>3</sub> will dissolve? Is the solution saturated, unsaturated, or supersaturated? 10 g of KClO<sub>3</sub> will dissolve, but 5 g will not, despite thorough stirring. The solution is saturated.

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SOLUTIONS 105

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## Chapter 3 Mixed Review Answers

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Interactive Reader

ChapTEr 1 Matter and

Change 2 1 2 3

Chemistry Is a Physical

Science 3 Matter and

Its Properties 6

Elements 16 ... two

chemicals are mixed.

Sometimes, scientists

do basic research

simply to satisfy their

curiosity about a

chemical and its

qualities. Sometimes

**Louisiana Interactive**

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## Chapter 3 Mixed Review Answers

### **Modern Chemistry Chapter 9 Stoichiometry Mixed Review Answers**

Report the number of significant figures in each of the following values: 3 a. 0.002 37 g 2 d. 64 mL 4 b. 0.002 037 g 2 e. 1.3 102 cm 3 c. 350. J 3 f. 1.30 102 cm 2. Write the value of the following operations using

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scientific notation. 103  
106 101 a. 102 8 103 4  
102 b. 5 2 10 4 4.3 10  
c. 3 103 4.0 104 3.

## **Chapter 2 Mixed Study Guide**

CHAPTER 2 REVIEW  
Measurements and  
Calculations MIXED  
REVIEW SHORT

ANSWER Answer the  
following questions in  
the space provided. 1.  
Match the description  
on the right to the  
most appropriate

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Review Answers

quantity on the left. d 2  
m<sup>3</sup> (a) mass of a small  
paper clip a 0.5 g (b)  
length of a small paper  
clip f 0.5 kg (c) length  
of a stretch limousine

## **2 Measurements and Calculations**

CHAPTER 1 REVIEW

Matter and Change

MIXED REVIEW SHORT

ANSWER Answer the  
following questions in  
the space provided. 1.  
Classify each of the  
following as a

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## Chapter 3 Mixed Review Answers

homogeneous or  
heterogeneous  
substance.

homogeneous a. sugar

homogeneous d.

plastic wrap

homogeneous b. iron

filings heterogeneous

e. cement sidewalk

heterogeneous c.

granola bar 2. For each

type of investigation,

select the most

appropriate ...

## **1 Matter and Change**

### **- HUBBARD'S**

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## Chapter 9 Mixed **CHEMISTRY**

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## **Modern Chemistry Chapter 9 Mixed Review Stoichiometry Answers**

3H<sub>4</sub>; 3 mol CO 2:1 mol  
C 3H<sub>4</sub>; 2 mol H<sub>2</sub>O:1  
mol C 3H<sub>4</sub>; 3 mol CO  
2:4 mol O<sub>2</sub>; 2 mol H<sub>2</sub>O:  
4 mol O<sub>2</sub>; 2 mol H<sub>2</sub>O:  
3 mol CO<sub>2</sub>; or their  
reciprocals b. Suppose  
that  $x$  moles of water

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## Chapter 3 Mixed Review Answers

form in the above reaction. The other three mole quantities (not in order) are  $2x$ ,  $1.5x$ , and  $0.5x$ . Match these quantities to their respective components in the equation ...

### **mc06se cFMsr i-vi**

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CHAPTER 9 REVIEW

Stoichiometry SECTION

3 PROBLEMS Write the

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Review Answers  
answer on the line to  
the left. Show all your  
work in the space  
provided. 1. 88% The  
actual yield of a  
reaction is 22 g and  
the theoretical yield is  
25 g. Calculate the  
percentage yield.

Copyright code: d41d8  
cd98f00b204e9800998  
ecf8427e.

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