

Chapter 9 Review Stoichiometry Answers Section 1

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Chapter 9 Review Stoichiometry Answers

CHAPTER 9 REVIEW Stoichiometry SHORT ANSWER Answer the following questions in the space provided. 1. Given the following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ 4 a. What is the value of the coefficient x in this equation? 40.07 g/mol b. What is the molar mass of C_3H_4 ? 2.1 mol H₂O c. What is the mole ratio of O₂ to H

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CHAPTER 9 REVIEW Stoichiometry

CHAPTER 9 REVIEW Stoichiometry MIXED REVIEW. SHORT ANSWER Answer the following questions in the space provided. 1. Given the following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ a. What is the value of the coefficient . x. in this equation? b. What is the molar mass of C_3H_4 ? c. How many moles are in an 8.0 g sample of C_3H_4 ? 2. a. What is meant by . ideal conditions

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CHAPTER 9 REVIEW. Stoichiometry. SECTION 9.2. PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. The following equation represents a laboratory preparation for oxygen gas:

CHAPTER 9 REVIEW

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Stoichiometry b. Theoretically, how many moles of NH₃ will be produced? PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N₂ are mixed with 12.0 mol of H₂ according to the ...

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Modern Chemistry Chapter 9 Mixed Review Stoichiometry Answers

Composition stoichiometry deals with the mass relationships of elements in compounds. Reaction stoichiometry involves the mass relationships between reactants and products in a chemical reaction. Reaction stoichiometry, the subject of this chapter, is based on chemical equations and the law of conservation of mass. All reaction stoichiometry

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CHAPTER 9 REVIEW Stoichiometry SECTION 2 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 4.5 mol The following equation represents a laboratory preparation for oxygen gas: $2KClO_3(s) \rightarrow 2KCl(s) + 3O_2(g)$ How many moles of O₂ form if 3.0 mol of $KClO_3$ are totally consumed?

Modern Chemistry Chapter 9 Review Stoichiometry Section 2 ...

CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left Show all your work in the space provided 1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g Calculate the percentage yield 2 60 mol of N₂ are mixed with 120 mol of H₂ according to the following equation: $N_2(g) + 3H_2(g)$

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